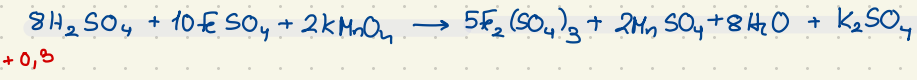
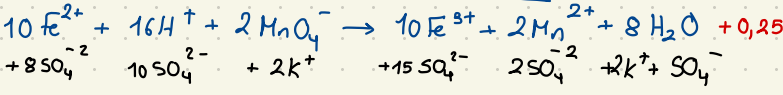
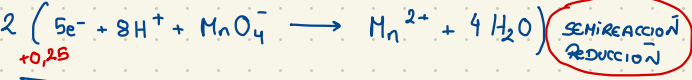
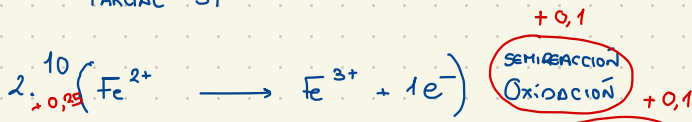


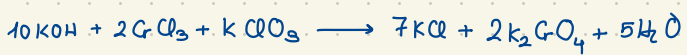
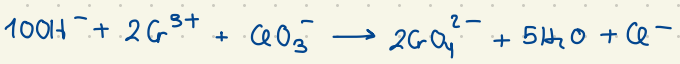
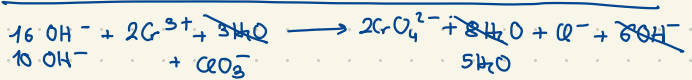
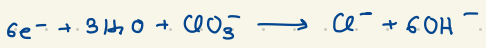
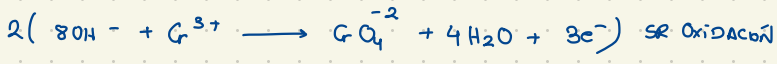
PARCIAL 3T



$$b) 7g \text{KMnO}_4 \cdot \frac{1 \text{ mol KMnO}_4}{188,04 \text{ g KMnO}_4} \cdot \frac{2 \text{ mol MnSO}_4}{2 \text{ mol MnSO}_4} \cdot \frac{151,7 \text{ g MnSO}_4}{1 \text{ mol MnSO}_4} \cdot \frac{75 \text{ g reales}}{100 \text{ g teorico}} = 5,02 \text{ g MnSO}_4$$

4. REGULADORA Ácido débil + SAL CONJUGADA  
 BASE ' + SAL CONJUGADA  
 FUNCIÓN: RESPONDER CAMBIOS PH

$$pH = 9 \Rightarrow \text{BASE} + \text{SAL}$$



3.

$$d = \frac{m}{V} \Rightarrow m = 1,2 \cdot 15 = 18 \text{ g disolución}$$

$$18 \text{ g disolución} \cdot \frac{37 \text{ g HCl}}{100 \text{ g disolución}} \cdot \frac{1 \text{ mol HCl}}{36,5 \text{ g HCl}} = 0,18 \text{ mol HCl}$$

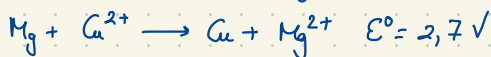
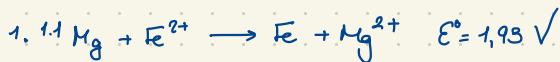
$$M = \frac{n}{V} = \frac{0,18}{0,15} = 1,2 \text{ M de HCl}$$



$$1,2 \frac{\text{mol}}{\text{L}} \cdot 0,02 \text{ L} \cdot \frac{1 \text{ mol NaOH}}{1 \text{ mol HCl}} \cdot \frac{1 \text{ L}}{0,5 \text{ mol NaOH}} = 0,048 \text{ L}$$

### 3.2. PROCEDIMIENTO

3.2. AZUL BROMOTIMOL  $\Rightarrow$  CAMBIO pH = 7



0,25 x 3 PREGUNDA PAR

0,25 JUSTIFICAR

### 1.2. ESQUEMA

$$1.3. \Delta G = -nFE \Rightarrow \Delta G = -2 \cdot 96500 \cdot 1,93 = -3,78 \cdot 10^5 \text{ J}$$

ESPONTÁNEA

