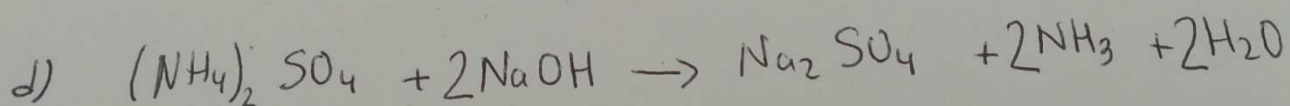
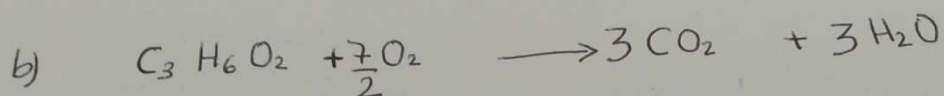
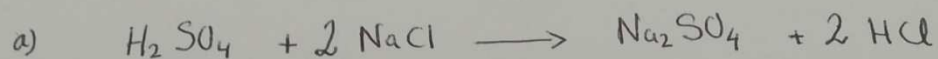


REACCIÓN QUÍMICAS

Resolución dos exercicios de reaccións químicas, poden facerse de outras formas diferentes e tamén poderían estar ben, esta só son unha mostra de como resolvelo. Se ao repasalos atopádesvos con calquera dúbida preguntádeme.

EXERCICIO 1

1.-



SHOT ON REDMI 7
AI DUAL CAMERA

EXERCICIO 2

2.

25 L O_2 en condições normais (1 mol gas em c.n. = 22,4 L)

$$25 \text{ L } O_2 \text{ em c.n. } \frac{1 \text{ mol c.n.}}{22,4 \text{ L}} = \boxed{1,12 \text{ mol } O_2}$$

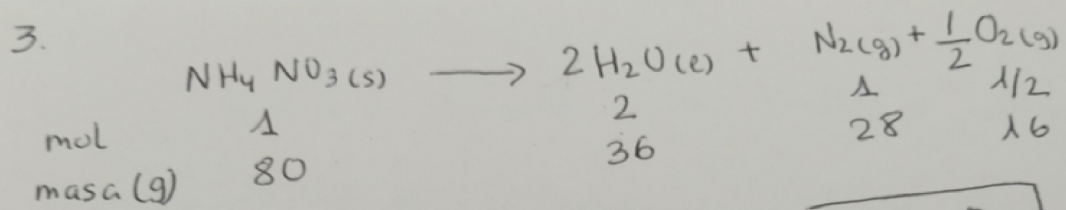
$$M_m(O_2) = 16 \cdot 2 = 32 \text{ g/mol}$$

$$1,12 \text{ mol } O_2 \cdot \frac{32 \text{ g}}{1 \text{ mol } O_2} = \boxed{35,84 \text{ g } O_2}$$

$$1,12 \text{ mol } O_2 \cdot \frac{6,02 \cdot 10^{23} \text{ moléculas}}{1 \text{ mol } O_2} = \boxed{6,74 \cdot 10^{23} \text{ moléculas}}$$

$$6,74 \cdot 10^{23} \text{ moléculas de } O_2 \cdot \frac{2 \text{ átomos O}}{1 \text{ molécula } O_2} = \boxed{1,348 \cdot 10^{24} \text{ átomos O}}$$

EXERCICIO 3



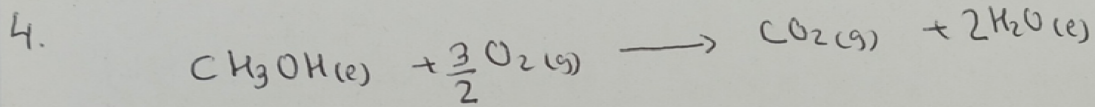
a) $100 \text{ g } \text{NH}_4\text{NO}_3 \cdot \frac{16 \text{ g } \text{O}_2}{80 \text{ g } \text{NH}_4\text{NO}_3} = \boxed{20 \text{ g } \text{O}_2}$

b) $20 \text{ g } \text{O}_2 \cdot \frac{1 \text{ mol } \text{O}_2}{32 \text{ g } \text{O}_2} = 0'625 \text{ mol } \text{O}_2$

$0'625 \text{ mol } \text{O}_2 \cdot \frac{22'4 \text{ L } \text{O}_2}{1 \text{ mol } \text{O}_2} = \boxed{14 \text{ L } \text{O}_2}$

EXERCICIO 4

$$0'625 \text{ mol } O_2 \cdot \frac{22'4 \text{ L } O_2}{1 \text{ mol } O_2} = \underline{14 \text{ L } O_2}$$



$$20 \text{ L } O_2 \text{ em c.n.} \cdot \frac{1 \text{ mol}}{22'4 \text{ L}} = 0'89 \text{ mol } O_2$$

a) $0'89 \text{ mol } O_2 \cdot \frac{1 \text{ mol } CH_3OH}{3/2 \text{ mol } O_2} = 0'60 \text{ mol } CH_3OH$

$$0'60 \text{ mol } CH_3OH \cdot \frac{32 \text{ g } CH_3OH}{1 \text{ mol } CH_3OH} = 19'2 \text{ g } CH_3OH$$

$$d = \frac{m}{V} \longrightarrow V = \frac{m}{d} = \frac{19'2}{0'792} = \underline{24'2 \text{ mL}}$$

b) $P \text{ e } T \text{ constantes}$

$$20 \text{ L } O_2 \cdot \frac{1 \text{ L } CO_2}{3/2 \text{ L } O_2} = \underline{13'3 \text{ L } CO_2}$$