

Ajusta la siguientes reacciones químicas:

1. $\text{H}_2 + \text{O}_2 \rightleftharpoons \text{H}_2\text{O}$
2. $\text{N}_2 + \text{H}_2 \rightleftharpoons \text{NH}_3$
3. $\text{H}_2\text{O} + \text{Na} \rightleftharpoons \text{Na(OH)} + \text{H}_2$
4. $\text{KClO}_3 \rightleftharpoons \text{KCl} + \text{O}_2$
5. $\text{BaO}_2 + \text{HCl} \rightleftharpoons \text{BaCl}_2 + \text{H}_2\text{O}_2$
6. $\text{H}_2\text{SO}_4 + \text{NaCl} \rightleftharpoons \text{Na}_2\text{SO}_4 + \text{HCl}$
7. $\text{FeS}_2 \rightleftharpoons \text{Fe}_3\text{S}_4 + \text{S}_2$
8. $\text{H}_2\text{SO}_4 + \text{C} \rightleftharpoons \text{H}_2\text{O} + \text{SO}_2 + \text{CO}_2$
9. $\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
10. $\text{NaCl} \rightleftharpoons \text{Na} + \text{Cl}_2$
11. $\text{HCl} + \text{MnO}_2 \rightleftharpoons \text{MnCl}_2 + \text{H}_2\text{O} + \text{Cl}_2$
12. $\text{K}_2\text{CO}_3 + \text{C} \rightleftharpoons \text{CO} + \text{K}$
13. $\text{Ag}_2\text{SO}_4 + \text{NaCl} \rightleftharpoons \text{Na}_2\text{SO}_4 + \text{AgCl}$
14. $\text{NaNO}_3 + \text{KCl} \rightleftharpoons \text{NaCl} + \text{KNO}_3$
15. $\text{Fe}_2\text{O}_3 + \text{CO} \rightleftharpoons \text{CO}_2 + \text{Fe}$
16. $\text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{NaHCO}_3$
17. $\text{FeS}_2 + \text{O}_2 \rightleftharpoons \text{Fe}_2\text{O}_3 + \text{SO}_2$
18. $\text{Cr}_2\text{O}_3 + \text{Al} \rightleftharpoons \text{Al}_2\text{O}_3 + \text{Cr}$
19. $\text{Ag} + \text{HNO}_3 \rightleftharpoons \text{NO} + \text{H}_2\text{O} + \text{AgNO}_3$
20. $\text{CuFeS}_2 + \text{O}_2 \rightleftharpoons \text{SO}_2 + \text{CuO} + \text{FeO}$